

Reg. No.



MANIPAL INSTITUTE OF TECHNOLOGY

MANIPAL

A Constituent Institution of Manipal University

II SEMESTER M.TECH.

END SEMESTER EXAMINATIONS, April 2017

SUBJECT: **FUEL CELL & HYDROGEN ENERGY [CHE5283]**

REVISED CREDIT SYSTEM

Time: 3 Hours

MAX. MARKS: 100

Instructions to Candidates:

- ❖ Answer **ALL** the questions.
- ❖ Missing data may be suitable assumed.

1A.	Define activation over potential and derive Tafel equation from butler volmer equation.	6																
1B.	Briefly describe about Electrochemical Impedance Spectroscopy? How do we represent EIS?	6																
1C.	Clearly explain the 1D fuel cell model on the basis of flux balance with the help of a neat diagram.	8																
2A.	Define Fuel Cell. Draw a neat diagram of a SOFC and Explain the half-cell reactions that take place in a fuel cell when transport ions are O^{2-} and CO_3^{-2} .	7																
2B.	Describe clearly about methods of H2 storage.	8																
2C.	<p>Calculate the metal dispersion and active metal area for 0.5 wt % Platinum on Alumina catalysts by CO pulse chemisorption. The areas of the pulses are given in the table. Additional given data are :</p> <p>Sample amount: 1gm ; Analysis gas: 10%CO/He; Carrier gas: He ; Temperature: 25°C ; Pulse volume: 0.5cm³ ; stoichiometry factor of CO on Pt = 1.</p> <table><tr><td>Peak No.</td><td>1</td><td>2</td><td>3</td><td>4</td><td>5</td><td>6</td><td>7</td></tr><tr><td>Peak area</td><td>0.0</td><td>0.0</td><td>0.0003</td><td>0.001</td><td>0.005</td><td>0.007</td><td>0.007</td></tr></table>	Peak No.	1	2	3	4	5	6	7	Peak area	0.0	0.0	0.0003	0.001	0.005	0.007	0.007	5
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3A.	What is Bragg's law? Determine the working principle and the instrumentation of XRD with the help of a neat diagram.	6																
3B.	Explain any two Ex situ characterization techniques of a fuel cell with the help of neat diagrams.	8																
3C.	Briefly explain the S-I cycle. Give advantages and disadvantages of this method.	6																
4A.	Write about direct and indirect reforming of fuel cells. What is the role of anode in this reforming technique? What are the advantages of direct internal reforming technique?	8																
4B.	How does Ohmic polarization arise in a fuel cell? Explain. Describe the three ways of transport in the electrolyte.	7																

4C.	What is Steam Reforming? Determine both the reactions for SR. write any 3 advantages and disadvantages of reforming reactions.	5															
5A.	Explain the working of FT-IR technique with the help of a diagram.	5															
5B.	<p>Give any two advantages and disadvantages of fuel cell. At 25°C, determine how does the electrical energy (ΔG) and heat produced differ if the water (the product) is in liquid and gaseous states???</p> <table> <tr> <th></th><th>hf (kJ/mol)</th><th>sf (kJ/mol.K)</th></tr> <tr> <td>H₂</td><td>0</td><td>0.13066</td></tr> <tr> <td>O₂</td><td>0</td><td>0.20517</td></tr> <tr> <td>H₂O (l)</td><td>-286.02</td><td>0.06996</td></tr> <tr> <td>H₂O (g)</td><td>-241.98</td><td>0.18884</td></tr> </table>		hf (kJ/mol)	sf (kJ/mol.K)	H ₂	0	0.13066	O ₂	0	0.20517	H ₂ O (l)	-286.02	0.06996	H ₂ O (g)	-241.98	0.18884	7
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5C.	Why do we use Chemisorption technique in characterizing a catalyst for fuel cell? Briefly explain working of pulse chemisorption technique with the help of a neat diagram.	8															